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(71) Applicant: BP Chemicals Limited  
London EC2M 7BA (GB)

(72) Inventors:

- Ellis, Brian,  
BP Chemicals Limited  
Sunbury on Thames, Middlesex TW16 7LN (GB)
- Jones, Michael David,  
BP Chemicals Limited  
Sunbury on Thames, Middlesex TW16 7LN (GB)

(74) Representative:

Perkins, Nicholas David et al  
BP International Limited  
Patents and Agreements Division,  
Chertsey Road  
Sunbury-on-Thames, Middlesex TW16 7LN (GB)

(54) Mixed metal oxide catalyst and use thereof in oxidation reactions

(57) A catalyst composition and its use for the oxidation of ethane to ethylene and/or acetic acid and/or for the oxidation of ethylene to acetic acid which comprises in combination with oxygen the elements molybdenum, vanadium, niobium and gold in the absence of palladium according to the empirical formula :  $\text{Mo}_a\text{W}_b\text{Au}_c\text{V}_d\text{Nb}_e\text{Y}_f$  (I) wherein Y is one or more elements selected from the group consisting of: Cr, Mn, Ta, Ti, B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Ir, Cu, Ag, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Zr, Hf, Ni, P, Pb, Sb, Si, Sn, Tl, U, Re, Te, La and Pd; a, b, c, d, e and f represent the gram atom ratios of the elements such that :  $0 < a \leq 1$ ;  $0 \leq b < 1$  and  $a + b = 1$ ;  $10^{-5} < c \leq 0.02$ ;  $0 < d \leq 2$ ;  $0 < e \leq 1$ ; and  $0 \leq f \leq 2$ .

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## Description

[0001] The present invention relates to a catalyst for the oxidation of ethane to ethylene and/or acetic acid and/or for the oxidation of ethylene to acetic acid, and to a process for the production of acetic acid and/or ethylene utilising the aforesaid catalyst. Catalysts comprising molybdenum, vanadium and niobium in combination with oxygen for use in processes for the production of acetic acid by the oxidation of ethane and ethylene are known in the art from, for example, US-A-4,250,346.

[0002] US Patent No. 4,250,346 discloses the oxidative dehydrogenation of ethane to ethylene in a gas phase reaction at relatively high levels of conversion, selectivity and productivity at a temperature less than 500°C using as catalyst a composition comprising the elements molybdenum, X and Y in the ratio  $Mo_aX_bY_c$  wherein X is Cr, Mn, Nb, Ta, Ti, V and/or W, and preferably Mn, Nb, V and/or W; Y is Bi, Ce, Co, Cu, Fe, K, Mg, Ni, P, Pb, Sb, Si, Sn, Tl and/or U, and preferably Sb, Ce and/or U, a is 1, b is 0.05 to 1.0 and c is 0 to 2, and preferably 0.05 to 1.0, with the proviso that the total value of c for Co, Ni and/or Fe is less than 0.5.

[0003] WO 99/20592 to Celanese relates to a method of selectively producing acetic acid from ethane, ethylene or mixtures thereof and oxygen at high temperature in the presence of a catalyst having the formula  $Mo_aPd_bX_cY_d$  wherein X represents one or several of Cr, Mn, Nb, Ta, Ti, V, Te and W; Y represents one or several of B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Ir, Cu, Ag, Au, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Nb, Zr, Hf, Ni, P, Pb, Sb, Si, Sn, Tl and U and a=1, b=0.0001 to 0.01, c = 0.4 to 1 and d = 0.005 to 1.

[0004] German patent application DE 196 30 832 AI relates to a similar catalyst composition in which a = 1, b > 0, c > 0 and d = 0 to 2. Preferably, a = 1, b = 0.0001 to 0.5, c = 0.1 to 1.0 and d = 0 to 1.0.

[0005] The catalysts of both WO 99/20592 and DE 19630832 require the presence of palladium.

[0006] WO 98/47850 relates to a process for producing acetic acid from ethane, ethylene or mixtures thereof and a catalyst having the formula  $W_aX_bY_cZ_d$  in which X represents one or several of Pd, Pt, Ag and Au, Y represents one or several of V, Nb, Cr, Mn, Fe, Sn, Sb, Cu, Zn, U, Ni, and Bi and Z represents one or several of Li, Na, K, Rb, Cs, Be, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ru, Os, Co, Rh, Ir, B, Al, Ga, In, Tl, Si, Ge, Pb, P, As and Te, a = 1, b > 0, c > 0 and d is 0 to 2.

[0007] The catalyst of WO 98/47850 does not contain molybdenum.

[0008] WO 99/51339 relates to a catalyst composition for the selective oxidation of ethane and/or ethylene to acetic acid which composition comprises in combination with oxygen the elements  $Mo_aW_bAg_cIr_dX_eY_f$  wherein X is the elements Nb and V; Y is one or more elements selected from the group consisting of Cr, Mn, Ta, Ti, B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Cu, Au, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Zr, Hf, Ni, P, Pb, Sb, Si, Sn, Tl, U, Re and Pd; a, b, c, d, e and f represent the gram atom ratios of the elements such that  $0 < a \leq 1$ ,  $0 \leq b < 1$  and  $a + b = 1$ ;  $0 < (c + d) \leq 0.1$ ;  $0 < e \leq 2$ ; and  $0 \leq f \leq 2$ .

[0009] There remains a need for catalyst for the oxidation of ethane to ethylene and/or acetic acid and/or for the oxidation of ethylene to acetic acid and a process for the production of acetic acid and/or ethylene employing the catalyst.

[0010] We have found that oxidation catalysts comprising molybdenum, vanadium and niobium in combination with oxygen which employ gold as an essential component in the absence of palladium can fulfil the need for an oxidation catalyst and process employing same.

[0011] Accordingly, the present invention provides a catalyst composition for the oxidation of ethane to ethylene and/or acetic acid and/or for the oxidation of ethylene to acetic acid which composition comprises in combination with oxygen the elements molybdenum, vanadium, niobium and gold in the absence of palladium according to the empirical formula:  $Mo_aW_bAu_cV_dNb_eY_f$  (I) wherein Y is one or more elements selected from the group consisting of: Cr, Mn, Ta, Ti, B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Ir, Cu, Ag, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Zr, Hf, Ni, P, Pb, Sb, Si, Sn, Tl, U, Re, Te, La and Pd; a, b, c, d, e and f represent the gram atom ratios of the elements such that:

$$0 < a \leq 1; 0 \leq b < 1 \text{ and } a + b = 1;$$

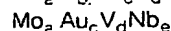
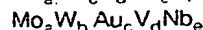
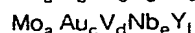
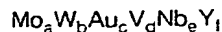
$$10^{-5} < c \leq 0.02;$$

$$0 < d \leq 2;$$

$$0 < e \leq 1; \text{ and}$$

$$0 \leq f \leq 2.$$

[0012] Catalysts embraced within the formula (I) include:-



[0013] Examples of suitable catalysts having the formula (I) include:  $\text{Mo}_{1.00}\text{V}_{0.25}\text{Nb}_{0.12}\text{Au}_{0.01}\text{O}_y$ ;  $\text{Mo}_{1.00}\text{V}_{0.213}\text{Nb}_{0.138}\text{Au}_{0.007}\text{O}_y$ ;  $\text{Mo}_{1.00}\text{V}_{0.232}\text{Nb}_{0.139}\text{Au}_{0.007}\text{O}_y$ ; and  $\text{Mo}_{1.000}\text{V}_{0.426}\text{Nb}_{0.115}\text{Au}_{0.0008}\text{O}_y$  wherein y is a number which satisfies the valencies of the elements in the composition for oxygen.

[0014] Preferably a > 0.01. Preferably, d > 0.1. Preferably, e > 0.01. Preferably, e ≤ 0.5. Preferably, f ≥ 0.01. Preferably, f ≤ 0.5.

[0015] Preferably, Y is selected from the group consisting of Bi, Ca, Ce, Cu, K, P, Sb, La and Te.

[0016] An advantage of catalyst compositions according to the present invention is that they can be more active and selective in converting ethane and/or ethylene to acetic acid and/or ethylene.

[0017] The catalyst compositions may be prepared by any of the methods conventionally employed for the preparation of catalysts. Suitably the catalyst may be prepared from a solution of soluble compounds and/or complexes and/or compounds of each of the metals. The solution is preferably an aqueous system having a pH in the range from 1 to 12, preferably from 2 to 8, at a temperature of from 20° to 100°C.

[0018] Generally, a mixture of compounds containing the elements is prepared by dissolving sufficient quantities of soluble compounds and dispersing any insoluble compounds so as to provide a desired gram-atom ratio of the elements in the catalyst composition. The catalyst composition may then be prepared by removing the solvent from the mixture. The catalyst may be calcined by heating to a temperature of from 200 to 550°C, suitably in air or oxygen, for a period of from 1 minute to 24 hours. Preferably, the air or oxygen is slowly flowing.

[0019] The catalyst may be used unsupported or supported. Suitable supports include silica, alumina, zirconia, titania, silicon carbide and mixtures of two or more thereof.

[0020] Further details of a suitable method for preparing a catalyst composition may be found in, for example, EP-A-0166438.

[0021] The catalyst may be used in the form of a fixed or a fluidised bed.

[0022] In another embodiment the present invention provides a process for the production of ethylene and/or acetic acid from a gaseous mixture comprising ethane and/or ethylene which process comprises contacting the gaseous mixture with a molecular oxygen-containing gas at elevated temperature in the presence of a catalyst composition as hereinbefore described.

[0023] Ethane is oxidised to ethylene and/or acetic acid and/or ethylene is oxidised to acetic acid. Preferably, ethane and optional ethylene is oxidised to a mixture comprising ethylene and acetic acid which may be used with or without the addition or removal of acetic acid and/or ethylene for the production of vinyl acetate by reaction with a molecular oxygen-containing gas in an integrated process.

[0024] The feed gas comprises ethane and/or ethylene, preferably ethane.

[0025] Ethane and/or ethylene may be used in substantially pure form or admixed with one or more of nitrogen, methane, carbon dioxide and water in the form of steam, which may be present in major amounts, for example greater than 5 volume percent or one or more of hydrogen, carbon monoxide,  $\text{C}_3/\text{C}_4$  alkenes and alkenes, which may be present in minor amounts, for example less than 5 volume percent.

[0026] The molecular oxygen-containing gas may be air or a gas richer or poorer in molecular oxygen than air, for example oxygen. A suitable gas may be, for example, oxygen diluted with a suitable diluent, for example nitrogen.

[0027] It is preferred to feed, in addition to ethane and/or ethylene and the molecular oxygen-containing gas, water (steam) because this can improve the selectivity to acetic acid.

[0028] The elevated temperature may suitably be in the range from 200 to 500°C, preferably from 200 to 400°C.

[0029] The pressure may suitably be atmospheric or superatmospheric, for example in the range from 1 to 50 bar, preferably from 1 to 30 bar.

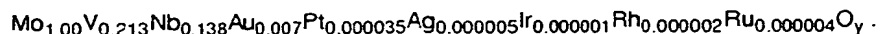
[0030] The catalyst composition is preferably calcined before use in the process of the invention. Calcination may suitably be achieved by heating at a temperature suitably in the range from 250 to 500°C in the presence of an oxygen-containing gas, for example air.

[0031] Operating conditions and other information applicable to the performance of the invention may be found in the aforesaid prior art, for example US Patent No. 4,250,346.

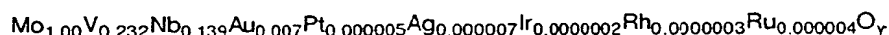
[0032] The process of the invention will now be further illustrated by reference to the following Examples.

CATALYST PREPARATIONExamples according to the present invention**Example 1 (Mo<sub>1.00</sub>V<sub>0.213</sub>Au<sub>0.007</sub>O<sub>y</sub>)**

[0033] A solution was prepared by dissolving 12.71 g ammonium molybdate, 2.11 g ammonium vanadate, 2.43 g niobium chloride, 0.36 g ammonium Au tetrachloride and 2.03 g oxalic acid in 200 ml water heated to 70°C for 15 minutes, followed by evaporation to dryness over 2 hours. The resulting catalyst cake was ground then calcined in static air in an oven at 350°C for 5 hours. The composition of the catalyst as analysed by XRF and found also to contain traces of platinum, silver, iridium, rhodium and ruthenium, represented as an empirical formula :

**Example 2 (Mo<sub>1.00</sub>V<sub>0.232</sub>Nb<sub>0.139</sub>Au<sub>0.007</sub>O<sub>y</sub>)**

[0034] A solution was prepared by dissolving 12.71 g ammonium molybdate, 2.11 g ammonium vanadate, 2.43 g niobium chloride, 0.36 g ammonium Au tetrachloride and 2.03 g oxalic acid in 200 ml water heated to 70°C for 15 minutes, followed by evaporation to dryness over 2 hours. The resulting catalyst cake was ground then calcined in static air in an oven at 350°C for 5 hours. The composition of the catalyst as analysed by XRF and found also to contain traces of platinum, silver, iridium, rhodium and ruthenium, represented as an empirical formula :

**Example 3 (Mo<sub>1.000</sub>V<sub>0.426</sub>Nb<sub>0.115</sub>Au<sub>0.0008</sub>O<sub>y</sub>)**

[0035] A solution was prepared by dissolving 12.92 g ammonium molybdate, 3.65 g ammonium vanadate, 2.27 g niobium chloride, 0.0214 g ammonium Au tetrachloride and 2.43 g oxalic acid in 200 ml water heated to 70°C for 15 minutes, followed by evaporation to dryness over 2 hours. The resulting catalyst cake was ground then calcined in static air in an oven at 400°C for 5 hours. The nominal composition of the catalyst was thus : Mo<sub>1.000</sub>V<sub>0.426</sub>Nb<sub>0.115</sub>Au<sub>0.0008</sub>O<sub>y</sub>

Examples not according to the invention**Comparative Example A. (Mo<sub>1.00</sub>V<sub>0.25</sub>Nb<sub>0.12</sub>O<sub>y</sub>)**

[0036] A solution A was prepared by dissolving 12.71 g ammonium molybdate in 50 ml water heated to 70°C. A second solution B was prepared by dissolving 2.11 g ammonium vanadate in 70 ml water heated to 70°C. Another solution C was prepared by dissolving 2.43 g niobium chloride and 2.02 g oxalic acid in 50 ml water heated to 70°C. Next, solution C was added to solution B and the resulting mixture heated at 70°C for 15 minutes. Solution A was then added and the final mixture heated at 70°C for 15 minutes, before evaporating the mixture to dryness over not more than 2 hours. The resulting catalyst cake was ground then calcined in static air in an oven at 350°C for 5 hours. The nominal composition of the catalyst was thus Mo<sub>1.00</sub>V<sub>0.25</sub>Nb<sub>0.12</sub>O<sub>y</sub>

**Comparative Example B (Mo<sub>1.00</sub>V<sub>0.25</sub>O<sub>y</sub>)**

[0037] As for comparative example A except that no solution C was prepared. The nominal composition of the catalyst was thus Mo<sub>1.00</sub>V<sub>0.25</sub>O<sub>y</sub>

**Comparative Example C (Mo<sub>1.00</sub>V<sub>0.25</sub>Nb<sub>0.12</sub>Pd<sub>0.01</sub>O<sub>y</sub>)**

[0038] As for comparative example A except that 0.23 g Pd acetate was added to solution A. The nominal composition of the catalyst was thus Mo<sub>1.00</sub>V<sub>0.25</sub>Nb<sub>0.12</sub>Pd<sub>0.01</sub>O<sub>y</sub>

**Comparative Example D (Mo<sub>1.00</sub>V<sub>0.25</sub>Nb<sub>0.12</sub>Ru<sub>0.01</sub>O<sub>y</sub>)**

[0039] As for comparative example A except that 0.36 g of ammonium Ru hexachloride was added to solution A. The nominal composition of the catalyst was thus Mo<sub>1.00</sub>V<sub>0.25</sub>Nb<sub>0.12</sub>Ru<sub>0.01</sub>O<sub>y</sub>

Comparative Example E ( $\text{Mo}_{1.00}\text{V}_{0.25}\text{Nb}_{0.12}\text{Rh}_{0.01}\text{O}_y$ )

[0040] As for comparative example A except that 0.15 g of Rh (III) hydroxide was added to solution A. The nominal composition of the catalyst was thus  $\text{Mo}_{1.00}\text{V}_{0.25}\text{Nb}_{0.12}\text{Rh}_{0.01}\text{O}_y$

CATALYST TEST PROCEDURE

[0041] Typically, 5 mls of catalyst was loaded into a fixed bed reactor made of Hastelloy grade C276 of dimensioning 12 mm internal diameter and length 40 cm. Glass beads were used to maintain the catalyst in position in the centre of the reactor. Above the catalyst, the glass beads thus acted as a mixing and pre-heating zone for gaseous and liquid reagents. The test apparatus was then pressure-tested at 21 bar with helium to check for leaks. Catalysts were then activated by heating to reaction pressure at 5° C/min in helium for 16 hours, to ensure full decomposition of catalyst precursors.

[0042] For Examples 1 and 2 and the Comparative Experiments, the required flows of ethane, 20 % oxygen in helium and water were then introduced to the reactor, to ensure an inlet composition of 42 % v/v ethane, 6.6 % oxygen, 25 % v/v helium and 26.4 % v/v water (as steam). The total feed flow rate was maintained at a level to ensure a feed GHSV of about 3000/h. In the case of Example 3, performance was measured with less water in the feed than Examples 1 and 2, and ethylene was added to the feed mixture, such that an inlet feed composition of 52 % v/v ethane, 5 % v/v ethylene, 6.6 % oxygen, 10 % v/v water and balance helium was used.

[0043] After equilibrating for 30 minutes, gas samples were taken from the outlet stream to calibrate a GC (model Unicam 4400) for ethane, oxygen and helium. Next, the setpoint temperature of the reactor was increased until 75 % - 100 % oxygen conversion was achieved, as indicated by the presence in the outlet stream of less than 2.2 % v/v oxygen.

[0044] Following an equilibration period of 30 minutes, catalysts were then evaluated under steady state conditions for a period of typically 4-5 hours. Exit gas volume was measured over the run period by a water-gas meter. Liquid products were collected and weighed after the run period. Composition of gas and liquid products was measured using GC analysis (Unicam 4400 and 4200 fitted with TCD and FID detectors respectively).

[0045] All feed and product flow rates and compositions were entered into a spreadsheet, and the following parameters calculated according to the formulae in which AcOH represents acetic acid :

$$\text{STY (space time yield) \%} = (\text{g AcOH}) / \text{litre catalyst bed} / \text{hour}$$

conversions :

[0046]

$$\text{of ethane} = (\text{inlet mol ethane} - \text{outlet mol ethane}) / \text{inlet mol ethane} * 100$$

$$\text{of oxygen} = (\text{inlet mol oxygen} - \text{outlet mol oxygen}) / \text{inlet mol oxygen} * 100$$

selectivities :

[0047]

$$\text{to AcOH (C-mol \%)} = (\text{outlet mol AcOH} * 2) / (\text{mol ethane converted} * 2) * 100$$

$$\text{to ethylene (C-mol \%)} = (\text{outlet mol ethylene} * 2) / (\text{mol ethane converted} * 2) * 100$$

$$\text{to CO (C-mol \%)} = (\text{outlet mol CO}) / (\text{mol ethane converted} * 2) * 100$$

$$\text{to CO}_2 \text{ (C-mol \%)} = (\text{outlet mol CO}_2) / (\text{mol ethane converted} * 2) * 100$$

$$\text{to CO}_x = \text{selectivity to CO (C-mol \%)} + \text{selectivity to CO}_2 \text{ (C-mol \%)}$$

Typically, mass balance and carbon balance for a reaction was found to be 100 +/- 5 %.

## RESULTS

[0048] The following tables 1 and 2 compare the catalyst performance of the above. Each catalyst was evaluated under standard conditions indicated in the tables, except for the temperature, which was varied so as to achieve 70-90 % oxygen conversion, in order to facilitate comparison.

Table 1

Experiment	Temp. of bed °C	Ethane convn %	Oxygen convn %	Sel. C <sub>2</sub> H <sub>4</sub> % C-mol	Sel CO <sub>x</sub> % C-mol	Sel. AcOH % C-mol	STY AcOH g/l-cat/h
A	326	3.2	82.4	41.9	16.2	41.3	82
B	314	4.8	79.4	15.1	67.0	17.8	20
C	253	3.9	76.8	0.0	36.9	62.9	82
D	293	5.0	71.8	43.7	16.5	39.8	70
E	304	5.7	87.5	43.1	15.5	40.3	89
Example 1	275	7.4	88.7	38.1	9.1	48.6	136
Example 1	285	9.2	100.0	37.8	8.5	51.5	177
Example 1	295	10.0	100.0	34.6	8.8	54.9	203
Example 2	303	7.1	73.3	38.1	6.3	54.6	142
Example 2	313	7.8	80.0	38.4	6.1	54.8	158
Example 2	323	8.8	92.7	40.8	7.2	51.1	165
Conditions : 42 % v/v ethane, 6.6 % v/v oxygen, 25 % v/v water, Balance He, GHSV=2970; 21 bar.							

Table 2

Experiment	Temp. of bed °C	Ethane convn %	Oxygen convn %	Sel. C <sub>2</sub> H <sub>4</sub> % C-mol	Sel CO <sub>x</sub> % C-mol	Sel. AcOH % C-mol	STY AcOH g/l-cat/h
Example 3	300	11.3	71.7	55.8	4.8	39.3	198
Feed : 52 % v/v ethane, 5 % v/v ethylene, 6.6 % v/v oxygen, 10 % v/v water, Balance He. Conditions : 9 bar, GHSV = 3200/h							

[0049] The results in the Tables illustrate clearly the promotional effect of gold on Mo-V-Nb based oxide catalysts. The results in the Tables also show that the performance of gold promoted catalysts according to the present invention is affected by the molar ratios of molybdenum, vanadium and niobium and by the gold content as well as being influenced by the reaction conditions.

## Claims

1. A catalyst composition for the oxidation of ethane to ethylene and/or acetic acid and/or for the oxidation of ethylene to acetic acid which composition comprises in combination with oxygen the elements molybdenum, vanadium, niobium and gold in the absence of palladium according to the empirical formula : Mo<sub>a</sub>W<sub>b</sub>Au<sub>c</sub>V<sub>d</sub>Nb<sub>e</sub>Y<sub>f</sub> (I) wherein Y is one or more elements selected from the group consisting of: Cr, Mn, Ta, Ti, B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Ir, Cu, Ag, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Zr, Hf, Ni, P, Pb, Sb, Si, Sn, Tl, U, Re, Te, La and Pd;

a, b, c, d, e and f represent the gram atom ratios of the elements such that :

$$0 < a \leq 1; 0 \leq b < 1 \text{ and } a + b = 1;$$

$$10^{-5} < c \leq 0.02;$$

$$0 < d \leq 2;$$

$$0 < e \leq 1; \text{ and}$$

$$0 \leq f \leq 2.$$

2. A catalyst composition as claimed in claim 1 wherein formula I is selected from the group consisting of  $\text{Mo}_a\text{W}_b\text{Au}_c\text{V}_d\text{Nb}_e\text{Y}_f$ ;  $\text{Mo}_a\text{Au}_c\text{V}_d\text{Nb}_e\text{Y}_f$ ;  $\text{Mo}_a\text{W}_b\text{Au}_c\text{V}_d\text{Nb}_e$  and  $\text{Mo}_a\text{Au}_c\text{V}_d\text{Nb}_e$ .

3. A catalyst composition as claimed in claim 1 or claim 2 wherein  $a > 0.01$ , and/or  $d > 0.1$ , and/or  $e > 0.01$  and/or  $e \leq 0.5$  and/or  $f \geq 0.01$  and/or  $f \leq 0.5$ .

4. A catalyst composition as claimed in any one of the preceding claims wherein Y is selected from the group consisting of Bi, Ca, Ce, Cu, K, P, Sb, La and Te.

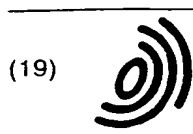
5. A catalyst composition as claimed in claim 1 having the formula selected from the group consisting of  $\text{Mo}_{1.00}\text{V}_{0.25}\text{Nb}_{0.12}\text{Au}_{0.01}\text{O}_y$ ;  $\text{Mo}_{1.00}\text{V}_{0.213}\text{Nb}_{0.138}\text{Au}_{0.007}\text{O}_y$ ;  $\text{Mo}_{1.00}\text{V}_{0.232}\text{Nb}_{0.139}\text{Au}_{0.007}\text{O}_y$ ; and  $\text{Mo}_{1.000}\text{V}_{0.426}\text{Nb}_{0.115}\text{Au}_{0.0008}\text{O}_y$  wherein y is a number which satisfies the valencies of the elements in the composition for oxygen.

6. A process for the oxidation of ethane to ethylene and/or acetic acid and/or for the oxidation of ethylene to acetic acid which process comprises contacting a gaseous mixture comprising ethane and/or ethylene with a molecular oxygen-containing gas at elevated temperature in the presence of a catalyst composition as claimed in any one of the preceding claims.

7. A process as claimed in claim 6 in which ethane and optional ethylene is oxidised to a mixture comprising ethylene and acetic acid.

8. A process as claimed in claim 7 in which the mixture comprising ethylene and acetic acid, with or without the addition or removal of acetic acid and/or ethylene, is further reacted with a molecular oxygen-containing gas to produce vinyl acetate.

9. A process as claimed in any one of claims 6 to 8 in which the catalyst is used in the form of a fluidised bed.



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(71) Applicant: **BP Chemicals Limited**  
London EC2M 7BA (GB)

(72) Inventors:  
• **Ellis, Brian, BP Chemicals Limited**  
Sunbury on Thames, Middlesex TW16 7LN (GB)  
• **Jones, Michael David, BP Chemicals Limited**  
Sunbury on Thames, Middlesex TW16 7LN (GB)

(74) Representative: **Perkins, Nicholas David et al**  
BP International Limited Patents and  
Agreements Division, Chertsey Road  
Sunbury-on-Thames, Middlesex TW16 7LN (GB)

(54) **Mixed metal oxide catalyst and use thereof in oxidation reactions**

(57) A catalyst composition and its use for the oxidation of ethane to ethylene and/or acetic acid and/or for the oxidation of ethylene to acetic acid which comprises in combination with oxygen the elements molybdenum, vanadium, niobium and gold in the absence of palladium according to the empirical formula:  $\text{Mo}_a\text{W}_b\text{Au}_c\text{V}_d\text{Nb}_e\text{Y}_f$  (I) wherein Y is one or more elements se-

lected from the group consisting of: Cr, Mn, Ta, Ti, B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Ir, Cu, Ag, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Zr, Hf, Ni, P, Pb, Sb, Si, Sn, Tl, U, Re, Te, La and Pd; a, b, c, d, e and f represent the gram atom ratios of the elements such that:  $0 < a \leq 1$ ;  $0 \leq b < 1$  and  $a + b = 1$ ;  $10^{-5} < c \leq 0.02$ ;  $0 < d \leq 2$ ;  $0 < e \leq 1$ ; and  $0 \leq f \leq 2$ .

EP 1 043 064 A3





European Patent  
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# PARTIAL EUROPEAN SEARCH REPORT

Application Number

which under Rule 45 of the European Patent Convention shall be considered, for the purposes of subsequent proceedings, as the European search report

EP 00 30 1108

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
X	EP 0 407 091 A (BP CHEM INT LTD) 9 January 1991 (1991-01-09) * page 2, line 50 - page 3, line 26 * * page 4, line 32 - page 5, line 3 * * page 5, line 21 - line 33 * * page 5, line 51 - page 6, line 14 * * claims * * column 2, line 31 - line 60 * * column 6, line 13 - line 22 *	1-9	B01J23/28 B01J23/20 B01J23/52 B01J23/68 C07C5/48 C07C11/04 C07C67/05 C07C69/01
X	US 4 250 346 A (THORSTEINSON ERLIND M ET AL) 10 February 1981 (1981-02-10) * column 2, line 14 - line 56 * * column 6, line 14 - line 22 * * examples 10,11,14,33-44,47-57; tables I-III * * claims *	1-9	
X	EP 0 294 845 A (UNION CARBIDE CORP) 14 December 1988 (1988-12-14) * page 3, line 24 - page 5, line 31 * * example 1 * * claims 1,2 *	1-9	
			TECHNICAL FIELDS SEARCHED (Int.Cl.7)
			B01J
<b>INCOMPLETE SEARCH</b> The Search Division considers that the present application, or one or more of its claims, does/do not comply with the EPC to such an extent that a meaningful search into the state of the art cannot be carried out, or can only be carried out partially, for these claims. Claims searched completely :  Claims searched incompletely :  Claims not searched :  Reason for the limitation of the search: see sheet C			
Place of search		Date of completion of the search	Examiner
MUNICH		19 October 2001	Gosselin, D
CATEGORY OF CITED DOCUMENTS X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons S : member of the same patent family, corresponding document			

EPO FORM 1303 (03.02) (Pct/CU)



European Patent  
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INCOMPLETE SEARCH  
SHEET C

Application Number  
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Claim(s) searched completely:

5

Claim(s) searched incompletely:

1-4,6-9

Reason for the limitation of the search:

Present claims 1 to 4 and 6 to 9 relate to an extremely large number of possible compounds and methods in which these compounds are involved. In fact, the claims contain so many options and a proviso that a lack of clarity (and conciseness) within the meaning of Article 84 EPC arises to such an extent as to render a meaningful search of the claims impossible. Consequently, the search has been carried out for those parts of the application which do appear to be clear (and concise), namely the compounds of claim 5 (of the examples of the application) and the process according to claims 6 to 9 as far as they are involved in it.

By searching the compounds of claim 5, closely related homologous compositions according to the second and fourth formulae of claim 2 (no tungsten oxide) were searched.

The disclaimer was not considered, since the wording of claim 1 is not consistent. Palladium is at the same time disclaimed as a whole and listed as element Y. Moreover, the reason of the disclaimer are not identified in the description.

The respective values of a, b, c, d, e and f in claims 1 and 3 were not necessarily considered.

In view of Example 1, it appears that values close to nil represent traces of compounds which are not voluntarily introduced in the composition. It has not been established that elements present in amount close to nil affect the catalytic activity of the overall composition. Such traces should also be present in the prior art by simply admixing the same main oxide precursors (Mo-V-Nb-Au).

Moreover, if a (0), c (lower value), d (0) and e (0) are close to nil, they should not be regarded as essential components of the catalyst but as traces as long as the lower limits from which they influence the overall catalytic activity have been established. On the limits any catalyst comprising tungsten oxide and only traces of the further elements identified in claim 1 should be regarded as novelty destroying for the subject-matter of claim 1.

The restrictions of claim 3 could not be considered. Due to an obscure wording, it was not possible to identify specific narrow ranges for a, d, c, d, e and f.

## PARTIAL EUROPEAN SEARCH REPORT

Application Number  
EP 00 30 1108

DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
X	DE 196 30 832 A (HOECHST AG) 5 February 1998 (1998-02-05) * page 2, line 44 - page 3, line 20 * * page 5, line 8 - line 27 * * claims *	1-9	
D,P, X	WO 99 51339 A (JONES MICHAEL DAVID ;KITCHEN SIMON JAMES (GB); COOK JOHN (GB); ELL) 14 October 1999 (1999-10-14) * page 4, line 8 - page 5, line 10 * * examples * * claims *	1-9	
A	US 3 567 772 A (YANAGITA MASAYA ET AL) 2 March 1971 (1971-03-02) * column 1, line 15 - line 32 * * example 11; table 8 *	1-5	TECHNICAL FIELDS SEARCHED (Int.Cl.7)

**ANNEX TO THE EUROPEAN SEARCH REPORT  
ON EUROPEAN PATENT APPLICATION NO.**

EP 00 30 1108

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report. The members are as contained in the European Patent Office EDP file on  
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19-10-2001

Patent document cited in search report		Publication date		Patent family member(s)	Publication date
EP 0407091	A	09-01-1991	AT	102913 T	15-04-1994
			AU	629004 B2	24-09-1992
			AU	5868190 A	10-01-1991
			BR	9003199 A	27-08-1991
			CA	2019509 A1	05-01-1991
			CN	1048539 A ,B	16-01-1991
			CN	1098664 A	15-02-1995
			DD	342531 B2	07-05-1992
			DE	8408211 U1	13-02-1986
			DE	69007359 D1	21-04-1994
			DE	69007359 T2	23-06-1994
			EP	0407091 A1	09-01-1991
			ES	2062372 T3	16-12-1994
			HU	54333 A2	28-02-1991
			JP	2837244 B2	14-12-1998
			JP	3044336 A	26-02-1991
			KR	146367 B1	17-08-1998
			NO	902935 A ,B,	07-01-1991
			PL	285929 A1	03-06-1991
			RU	2002730 C1	15-11-1993
			US	5260250 A	09-11-1993
			US	5210293 A	11-05-1993
<hr/>					
US 4250346	A	10-02-1981	NONE		
<hr/>					
EP 0294845	A	14-12-1988	US	5162578 A	10-11-1992
			AU	608615 B2	11-04-1991
			AU	1758688 A	15-12-1988
			CA	1305179 A1	14-07-1992
			DE	3870717 D1	11-06-1992
			EP	0294845 A1	14-12-1988
			JP	1085945 A	30-03-1989
			ZA	8804181 A	22-02-1989
<hr/>					
DE 19630832	A	05-02-1998	DE	19630832 A1	05-02-1998
			AU	716609 B2	02-03-2000
			AU	4010597 A	25-02-1998
			BG	103119 A	29-10-1999
			BR	9710906 A	17-08-1999
			CN	1231656 A	13-10-1999
			CZ	9900316 A3	14-07-1999
			WO	9805619 A1	12-02-1998
			EP	0915821 A1	19-05-1999
			JP	2000515158 T	14-11-2000
			NO	990363 A	26-01-1999
			PL	331374 A1	05-07-1999

EPO FORM P0409

For more details about this annex : see Official Journal of the European Patent Office, No. 12/82

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ON EUROPEAN PATENT APPLICATION NO.**

EP 00 30 1108

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The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

19-10-2001

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
DE 19630832	A		SI 9720051 A	31-10-1999
			TR 9900162 T2	22-03-1999
			TW 427971 B	01-04-2001
			US 6194610 B1	27-02-2001
WO 9951339	A	14-10-1999	AU 3160899 A	25-10-1999
			CN 1303320 T	11-07-2001
			EP 1069945 A1	24-01-2001
			WO 9951339 A1	14-10-1999
			NO 20004909 A	17-11-2000
			TR 200002817 T2	21-12-2000
US 3567772	A	02-03-1971	DE 1643087 A1	06-05-1971
			GB 1170851 A	19-11-1969